

Subject: Chemistry (AQA)

Pre-course Learning

Introduction

If you're keen to study this A-Level at The Quest Academy, below is some useful information about the course and some pre-learning that you can engage with over the summer. Upon commencing the course in September, you'll receive a printed Welcome Pack which will detail the order of the topics you'll be taught and your teachers, assessment dates and revision lists, how to keep organised for the course, how to revise for Chemistry and some other useful information and links.

Course Outline

The exam board for the course is Edexcel

Paper	Texts	% of A Level	Assessment Objectives
Paper 1 AS	written exam: 1 hour 30 minutes 80 marks 50% of the AS 65 marks of short and long answer questions 15 marks of multiple choice questions	50 % of AS	Relevant physical chemistry topics (sections 3.1.1 to 3.1.4, 3.1.6 and 3.1.7) Inorganic chemistry (section 3.2.1 to 3.2.3) Relevant practical skills
Paper 2 AS	written exam: 1 hour 30 minutes 80 marks 65 marks of short and long answer questions 15 marks of multiple choice questions	50% of AS	Relevant physical chemistry topics (sections 3.1.2 to 3.1.6) Organic chemistry (section 3.3.1 to 3.3.6) Relevant practical skills
Paper 1 A-Level	written exam: 2 hours 105 marks	35% of A Level	Relevant physical chemistry topics (sections 3.1.2 to 3.1.6 and 3.1.9) Organic chemistry (section 3.3) Relevant practical skills
Paper 2 A-Level	written exam: 2 hours	35% of A-Level	Relevant physical chemistry topics

	105 marks 105 marks of short and long answer questions		(sections 3.1.2 to 3.1.6 and 3.1.9) Organic chemistry (section 3.3) Relevant practical skills
Paper 3 A-Level	written exam: 2 hours 90 marks 40 marks of questions on practical techniques and data analysis 20 marks of questions testing across the specification 30 marks of multiple choice questions	30% of A-Level	Any content Any practical skills

Year 12

Autumn 1	Amount of substance, Atomic Structure
Autumn 2	Introduction to alkanes , alkenes, Bonding, Periodicity
Spring 1	Haloalkanes, Redox, Group 2 Group 7
Spring 2	Kinetics, Equilibria and Kc
Summer 1	Energetics
Summer 2	Kp

Year 13

Autumn 1	Isomerism, Aldehydes and ketones, Aromatics, Amines and Revision
Autumn 2	Carboxylic acids and their derivatives, rate equations, Polymers, Thermodynamics
Spring 1	Amino acids, Protein and DNA, Chromatography
Spring 2	Organic synthesis, NMR , Transition metals, Acids and bases
Summer 1	Electrochemistry
Summer 2	Revision and catch up of practicals

Tasks

Task 1: Bonding

Decide if the compounds below are ionically or covalently bonded together and why?

- a) Ammonia NH_3
- b) Zinc Oxide ZnO
- c) Methane CH_4
- d) Benzene C_6H_6
- e) Potassium Dichromate $\text{K}_2\text{Cr}_2\text{O}_7$

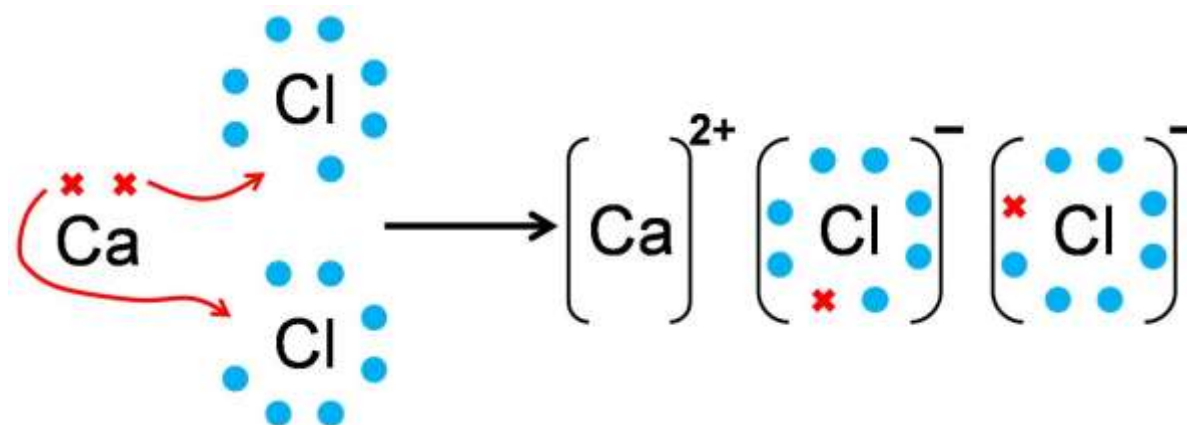
Ionic Bonding

This is an ELECTROSTATIC ATTRACTION between 2 oppositely charged species called IONS.

The compound is formed is neutral, which means it has no overall charge.

i.e. it has an equal amount of positive and negative charge from the different ions that are making it up.

How are IONS made?



This is seen by the diagram above:

METALS: (Calcium)	NON- METALS (Chlorine)
They form Positive ions as they lose their outer electrons to form a FULL OUTER SHELL.	They form NEGATIVE ions as they gain electrons to form a FULL OUTER SHELL.
Calcium 2 electrons in its outer shell as an element so LOSES 2 electrons to become a 2+ ion	Chlorine has 7 electrons in its outer shell so will GAIN 1 electron to become a 1- ion

Task 2: Formation of Ionic compounds

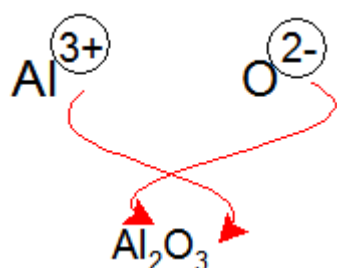
Draw out Atom and Ions for the following ionic compounds (like the calcium Oxide diagram above)

- 1) Aluminium Oxide
- 2) Lithium Oxide
- 3) Barium Nitride

Formula of Ionic compounds

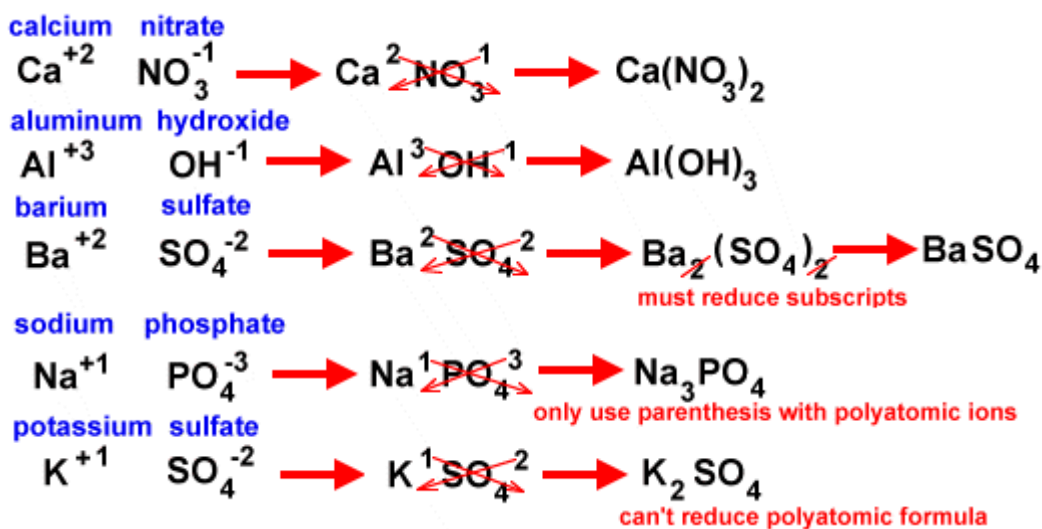
When we form an Ionic compound we have oppositely charged ions attracted together so that a neutral compound is formed.

This means there is a balance between the positive metals ions and negative non-metal ions.



You swap the NUMBERS of the charge over
If a 1 you ignore it
If get 2 numbers the same ignore them

Aluminium Oxide made from Aluminium ions and Oxide ions.



Other examples above (don't worry about the writing in red)

Task 3: Formula Interpretation

Using the table of common ions work out the formula of the following ionic compounds.

- 1) Silver chloride
- 2) Lithium sulphate
- 3) Ammonium Hydroxide
- 4) Potassium Dichromate
- 5) Iron (II) Nitrate

Formula interpretation

When we have calculated the formula of a compound it is important we can interpret the information about the number of atoms and types of elements in the compound.

e.g.

Calcium Carbonate

CaCO_3

1 Ca

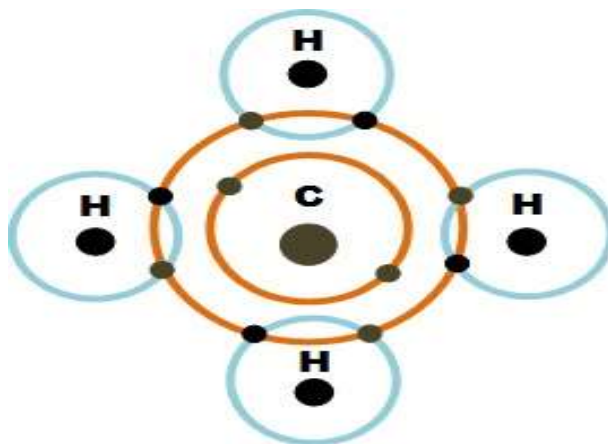
1 C

3 O

Task 4: Covalent Bonding

Look at the following compounds and work out the number and type of elements in the compound.

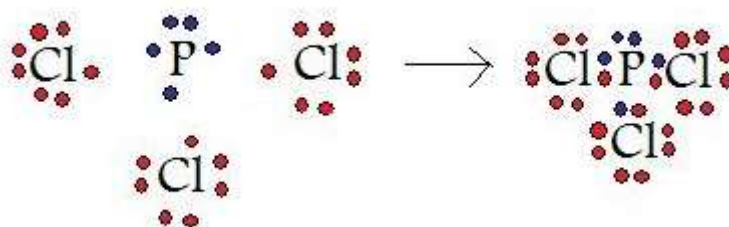
- 1) AgNO_3
- 2) PbCO_3
- 3) SnCl_2
- 4) Mg(OH)_2



Covalent bonding

The covalent bond is made up from non-metal atoms that want to bond together.

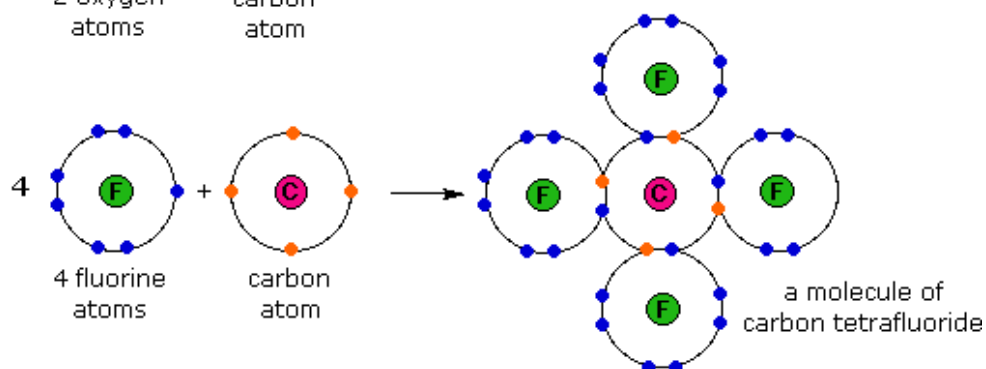
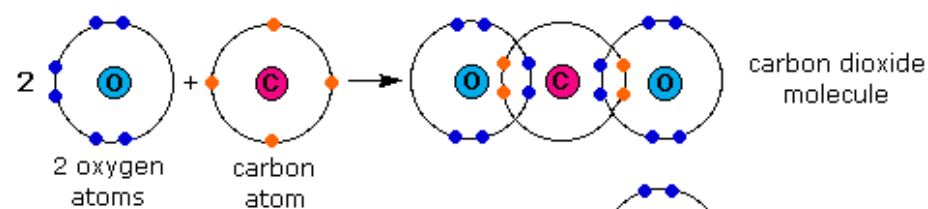
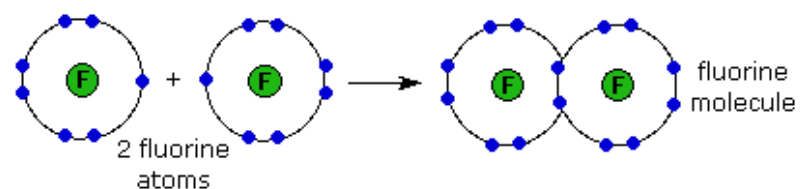
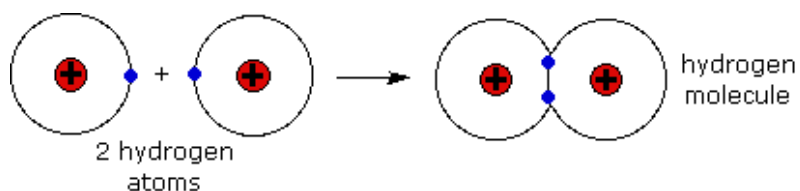
Covalent bonds are made from the atoms sharing their electrons to get a FULL OUTER SHELL.



The above example shows,

Phosphorus in group 5 with 5 outer electrons sharing 1 electron each with a chlorine atom which is in group 7.

Both the Phosphorus and Chlorine NOW have their FULL OUTER SHELL.



More examples

The example shows a series of covalently bonded molecules where the atoms have all got a FULL OUTER SHELL.

Please note

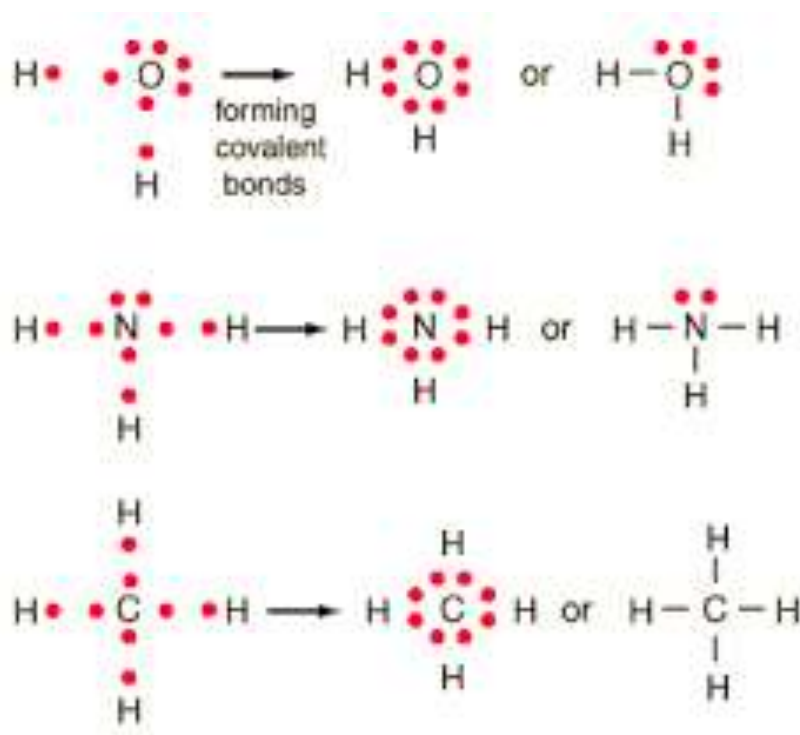
DOUBLE BOND on the CO_2 molecule .

The 4 SINGLE BONDS from the carbon attached to each individual F in the CF_4 molecule.

EXT Line diagrams

These are simpler versions of the shown DOT-CROSS diagrams where you show each bond (PAIR of ELECTRONS) as a line between the atoms in the molecule

e.g.



The extra pair of electrons that are not involved in the bonds are called LONE PAIR of electrons. These are shown by the pair of 'dots' around the central atom.

Task 5

Draw out the Dot/ Cross diagrams and Line diagram of the following molecules:

- 1) Ethane C_2H_6
- 2) Propene C_3H_6
- 3) Hydrogen Peroxide H_2O_2
- 4) Hydrogen Sulphide H_2S

Chapter 2

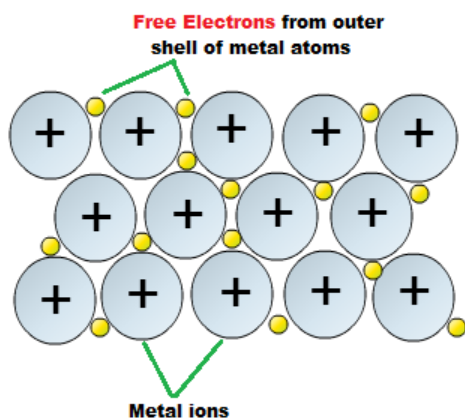
Structure

There are 4 main structures you need to be aware of

- 1) Metallic structure
- 2) Giant Ionic
- 3) Giant covalent / Macromolecular
- 4) Simple Molecular

1 Metallic

This occurs in metals.

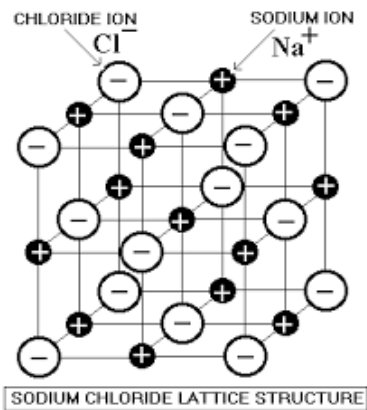


These are strongly bonded structures which have HIGH boiling and melting points.
They CAN conduct electricity due to the FREE ELECTRONS.

2 Giant Ionic

This occurs as a LATTICE of IONS electrostatically attached together with the positive ions being attracted to the negative ions.

It occurs in Ionically bonded compounds.



3 Giant covalent / Macromolecular

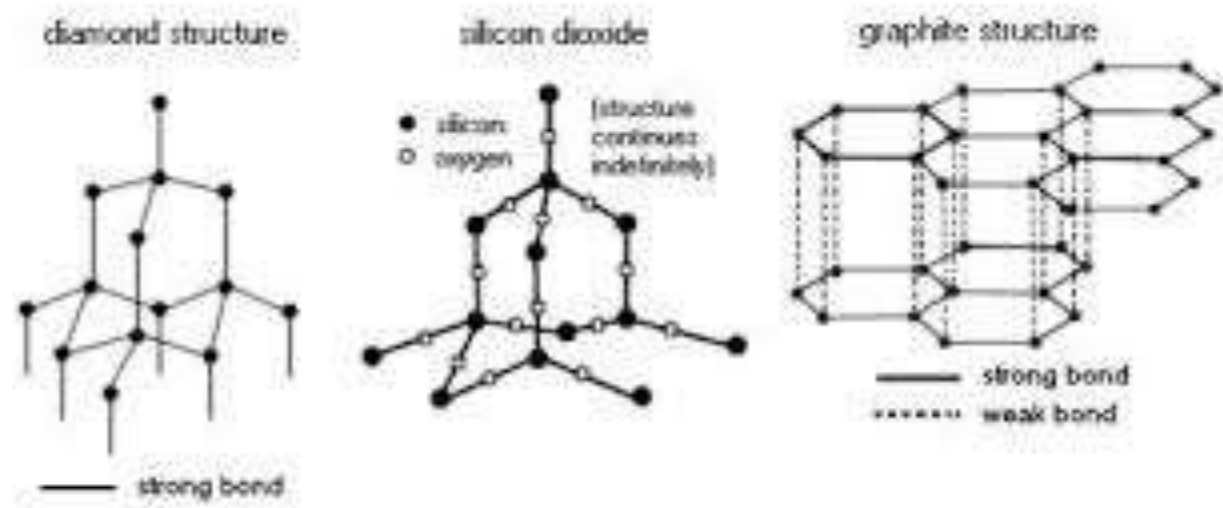
This occurs in a select number of covalently bonded compounds which have ALL their atoms covalently bonded together in a large structure.

Key examples are ALLOTROPES of carbon (look up what Allotrope means!) and silicon dioxide

Diamond

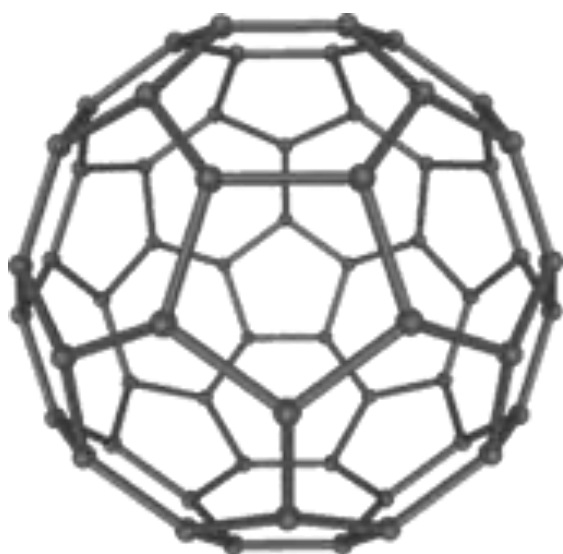
Silicon Dioxide

Graphite



EXT

Buckminster Fullerene



This is a C_{60} molecule in the shape of a football.
They were discovered in the UK in 1985 and the chemists involved won the Nobel prize in 1996.

4) Simple Molecular

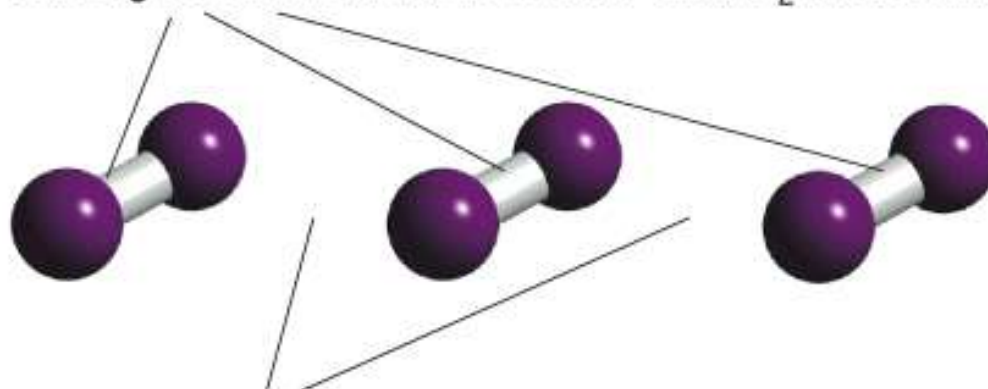
This occurs in covalently bonded molecules which have **STRONG** covalent bonds inside the molecules
But

Much weaker **INTERMOLECULAR** bonds between the molecules.

The three types of **INTERMOLECULAR** bond/ force are:

- Van Der Waals
- Permanent Dipole
- Hydrogen Bond

Strong covalent bonds within each I_2 molecule



Weak van der Waals' forces between I_2 molecules

Task 6

Research task

Find out what the trend in melting/ boiling point is for Na-Mg-Al (the metal in the third period)
Explain why there is this trend (linked to their structure)

<http://www.creative-chemistry.org.uk/alevel/module1/trends8.htm>

(basic source exemplar)

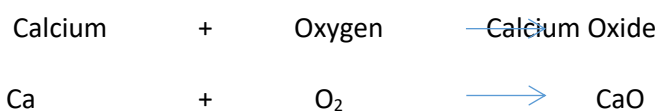
Chapter 3

Equations

We will be most interested in BALANCED symbol equations.

These show us exactly what elements are in the reactants and the products and we need the SAME amount on both sides of the equation.

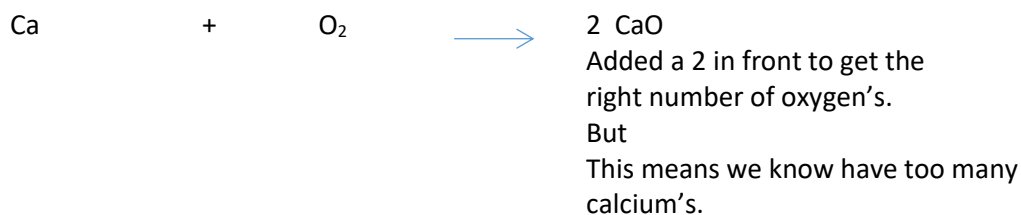
Example



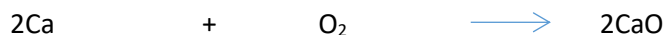
This is not balanced,

So we need to ADD large numbers in front of the formula given to balance it.

Firstly



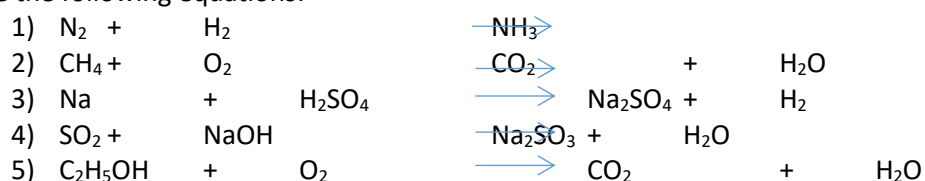
So we now need to add
2 on this side as well



It is now a Balanced equation.

Task 7

Balance the following equations:



State symbols

These are linked to the three states of matter

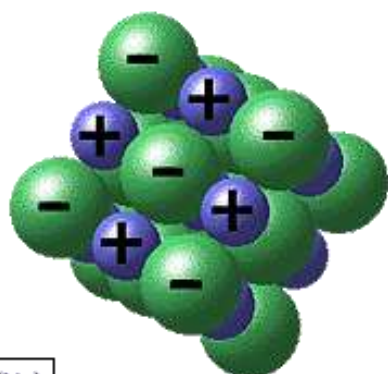
- Gas (g)
- Liquid (l)
- Solid (s)

Also we have (aq) for a solution.

EXT

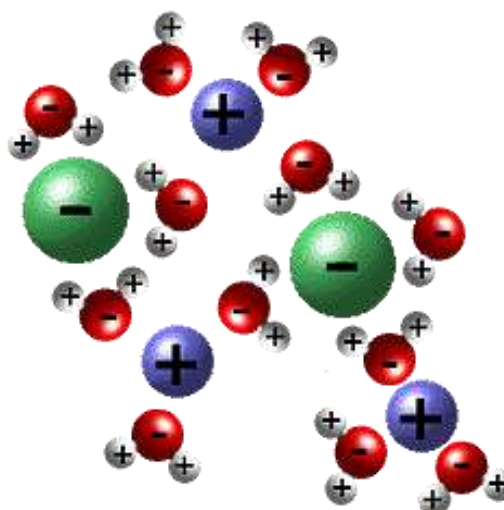
Ionic compounds in solutions

NaCl crystal structure



sodium (Na)
chlorine (Cl)

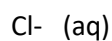
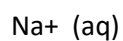
NaCl in water

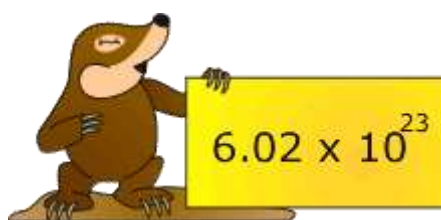


When we dissolve an ionic compound it is the separate ions in the compound being split apart and bonded to the water.



Is in fact





Chapter 4

Mole work.

In its most basic form the 'MOLE' is just a word used to describe a number.

e.g. Couple 2
 Dozen 12
 Mole 6.02×10^{23} (602000000000000000000000)

Why this large number?

It was found that this number of ATOMS of any element is equal to the MASS NUMBER of this element in grams.

e.g.
 6.02×10^{23} carbon atoms is equal to 12g
 6.02×10^{23} neon atoms is equal to 20g

This leads to the FIRST mole equation.

$$\text{Moles} = \frac{\text{Mass}}{\text{R.A.M}} \quad (\text{relative atomic mass})$$

e.g.
 How many moles are there in 24g of carbon?

$$\text{Moles} = \frac{\text{Mass}}{\text{R.A.M}}$$

$$\text{Moles} = \frac{24}{12}$$

$$\text{Moles} = 2 \text{ moles of carbon}$$

Task 8

Calculate the number of moles in the following elements?

- 1) 59 g of cobalt
- 2) 4.14 g of lead
- 3) 1.08g of gold

This can get increased very quickly to include compounds and not just elements.

In this we use a very similar Mole equation:

$$\text{Moles} = \frac{\text{Mass}}{\text{R.F.M}} \quad \text{This is the Relative formula mass}$$

e.g. H_2O

$\text{H} + \text{H} + \text{O}$

$$1 + 1 + 16 = 18$$

e.g.

How many moles are there in 88g of carbon dioxide?

$$\begin{aligned} \text{Moles} &= \frac{\text{Mass}}{\text{R.F.M}} \\ &= \frac{88}{44} \\ &= 2 \text{ mole} \end{aligned} \quad \begin{aligned} &\text{CO}_2 \\ &\text{C} + \text{O} + \text{O} \\ &12 + 16 + 16 = 44 \end{aligned}$$

NOTE- Good practice

It is always good practice to start with the equation in word form then put the numbers in from the questions

It is also good practice to show how you have worked out the RFM so if there is an error you can still get method marks.

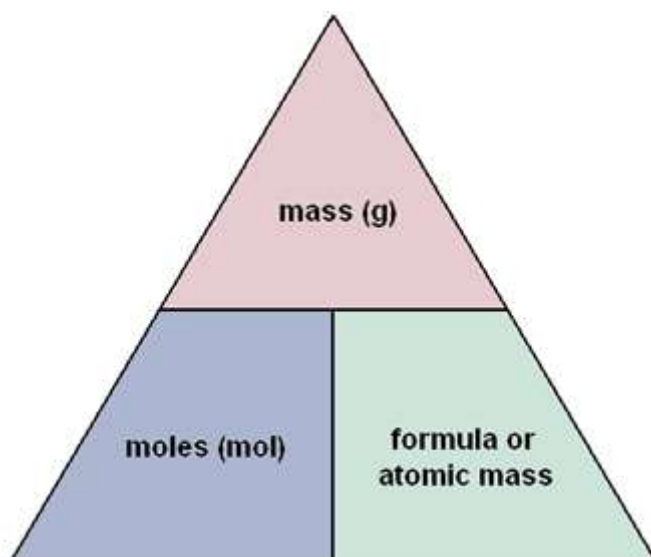
Task 9

How many moles are there in the following:

- 1) 62 g of sodium Oxide Na_2O
- 2) 174 g of lithium bromide LiBr
- 3) 3.2 g of oxygen
- 4) 1.24 g of Ammonia

Changing the equation

We can have this mole equation in a simple MAGIC TRIANGLE and easily change the aspect we are trying to work out.



So we may get asked to calculate the Mass or Relative formula mass.

Task 10

Calculate the :

- 1) Mass of 2 moles of calcium metal
- 2) 0.25 moles of lead carbonate PbCO_3
- 3) The formula mass of a compound which has 0.5 moles of mass 14g

EXT

Harder question

Task 11

250g of hydrated copper sulphate ($\text{CuSO}_4 \cdot x \text{H}_2\text{O}$) is collected and a student want to calculate the number of moles of water attached to the copper sulphate, the x value.

The student completely dried the copper sulphate and the new mass was found to be 160g

- a) Calculate the moles of copper sulphate
- b) Calculate the mass of lost water
- c) Calculate the number of moles of lost water
- d) Therefore deduce the formula of the hydrated copper sulphate.

Moles and solution

When we dissolve a solid in water we create a solution.

We use a different mole equation to calculate the moles in the solutions we create.

$$\text{Moles} = \frac{\text{Conc} \times \text{Vol}}{1000}$$

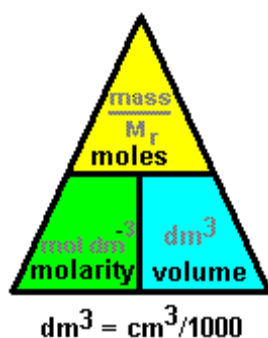
Molarity / M
Mol/dm³

ml or cm³

e.g.

How many moles are there in 250 cm³ of 0.1 M Hydrochloric acid ?

$$\begin{aligned} \text{Moles} &= \frac{\text{Conc} \times \text{Vol}}{1000} \\ &= \frac{0.1 \times 250}{1000} \\ &= 0.025 \text{ Moles} \end{aligned}$$



This equation can again be moved around if you have to calculate the concentration using the moles and volume.

Task 12

- 1) Calculate the moles in 40 ml of 5M of sodium hydroxide solution
- 2) What is the concentration when you dissolve 2 moles of acid in 100ml of water
- 3) How many moles are there in 500ml of 0.1 mol/dm³ of salt solution
- 4) What is the concentration of 0.25 moles of alkali in 25 ml

EXT

Combining our work

We often need to combine this work on moles and work out the mass of a solid we need to make up a set concentration of a solution.

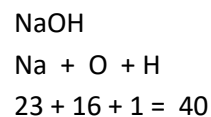
I.e. we want to make 100ml volume of a 0.5 M solution of sodium Hydroxide, how much mass do we need to dissolve?

- 1) How many moles are in this solution,

$$\begin{aligned}
 \text{Moles} &= \frac{\text{Conc} \times \text{Vol}}{1000} \\
 &= \frac{0.5 \text{ M} \times 100\text{ml}}{1000} \\
 &= \underline{0.05} \text{ Moles of sodium hydroxide in solution}
 \end{aligned}$$

- 2) What mass do we need for that many moles,

$$\begin{aligned}
 \text{Mass} &= \text{moles} \times \text{RFM} \\
 &= 0.05 \times 40 \\
 &= \underline{2 \text{ g}}
 \end{aligned}$$



So we will need to dissolve 2 g in the 100ml to make the required solution concentration of 0.5M.

Task 13

- 1) How many grams of potassium oxide (K_2O) are needed to make 100ml of a 0.5M solution ?
- 2) What is the concentration of a solution when we dissolve 730g of hydrochloric acid in 350 cm^3 ?
- 3) What is the mass of calcium oxide, CaO needed to make a 250 ml volume of 0.5 M solution?

NOTE- HINT

Keep looking carefully at the units

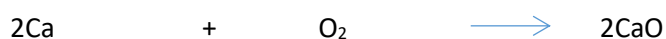
MI= cm^3 for volume mol/dm^3 = Molarity =M for concentration

Molar Ratio

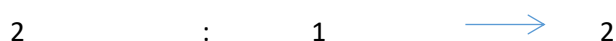
This is the link between the balanced symbol equations and the amount of moles of each substance in the reaction.

Simply it is the ratio of the numbers in front of the compounds in the balanced symbol equation.

e.g.



In this equation the Molar ratio is:

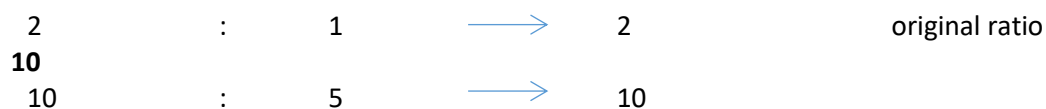


Means:

2 moles of calcium will react with 1 mole of oxygen and we will make 2 moles of the calcium oxide.

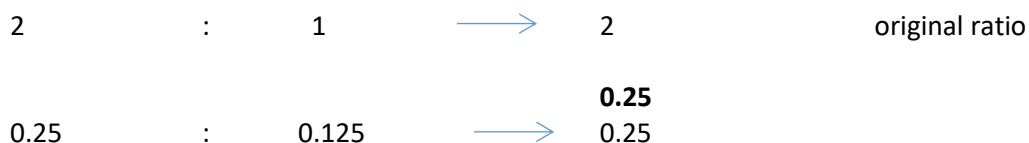
As it is a ratio these numbers can be varied,

So if we actually had **10** moles of the calcium?



So 10 moles of the calcium would react with 5 moles of the oxygen and form 10 moles of the calcium oxide

Or if we wanted to make **0.25** moles of the calcium oxide



We would need 0.25 moles of the CaO

Final mole equation work

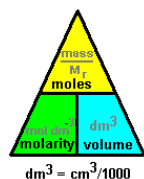
We are often asked to calculate how much we will produce in a reaction from a certain starting amount of reactants, or how much reactants we will need to make a set amount of products.

We put together the :

- Molar ratio work with the balanced equation
- The different moles equations

NOTE

If it involves a SOLID it is ...



If it involves a solution it is ..

e.g.

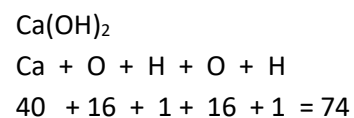
Calcium oxide reacts with water to form calcium hydroxide.



If I started with 28g of the calcium oxide what mass of calcium hydroxide would I make, and if it was in 100ml of water what would its concentration be



$$\begin{aligned} \text{Moles} &= \frac{\text{Mass}}{\text{RFM}} \\ &= \frac{28}{56} \\ &= 0.5 \text{ moles} \end{aligned}$$



$$\begin{array}{ccccc} 0.5 & & 0.5 & & 0.5 \quad \text{New molar ratio} \end{array}$$

$$\begin{aligned} \text{Mass} &= \text{Moles} \times \text{RFM} \\ &= 0.5 \times 74 \\ &= \underline{37\text{g}} \end{aligned}$$

And the solution concentration would be:

$$\begin{array}{l} 0.5 \text{ moles} \\ 100\text{ml} \end{array}$$

$$\text{Conc} = \frac{1000 \times \text{mole}}{\text{Vol}}$$

$$\text{Conc} = \frac{1000 \times 0.5}{100}$$

$$\text{Conc} = 5 \text{ mol/dm}^3$$

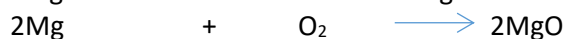
Task 14

- 1) Calcium cyanamide CaCN_2 reacts with water to form calcium carbonate and ammonia



What mass of calcium carbonate is formed if 20g of the CaCN_2 is reacted with excess water.

- 2) Magnesium burns in air to make magnesium oxide



What mass of magnesium would you need to create 0.8g of magnesium oxide powder.

- 3) Iron reacts with water to form iron oxide and hydrogen



If the student starts with 1.68g of iron and it undergoes a complete reaction

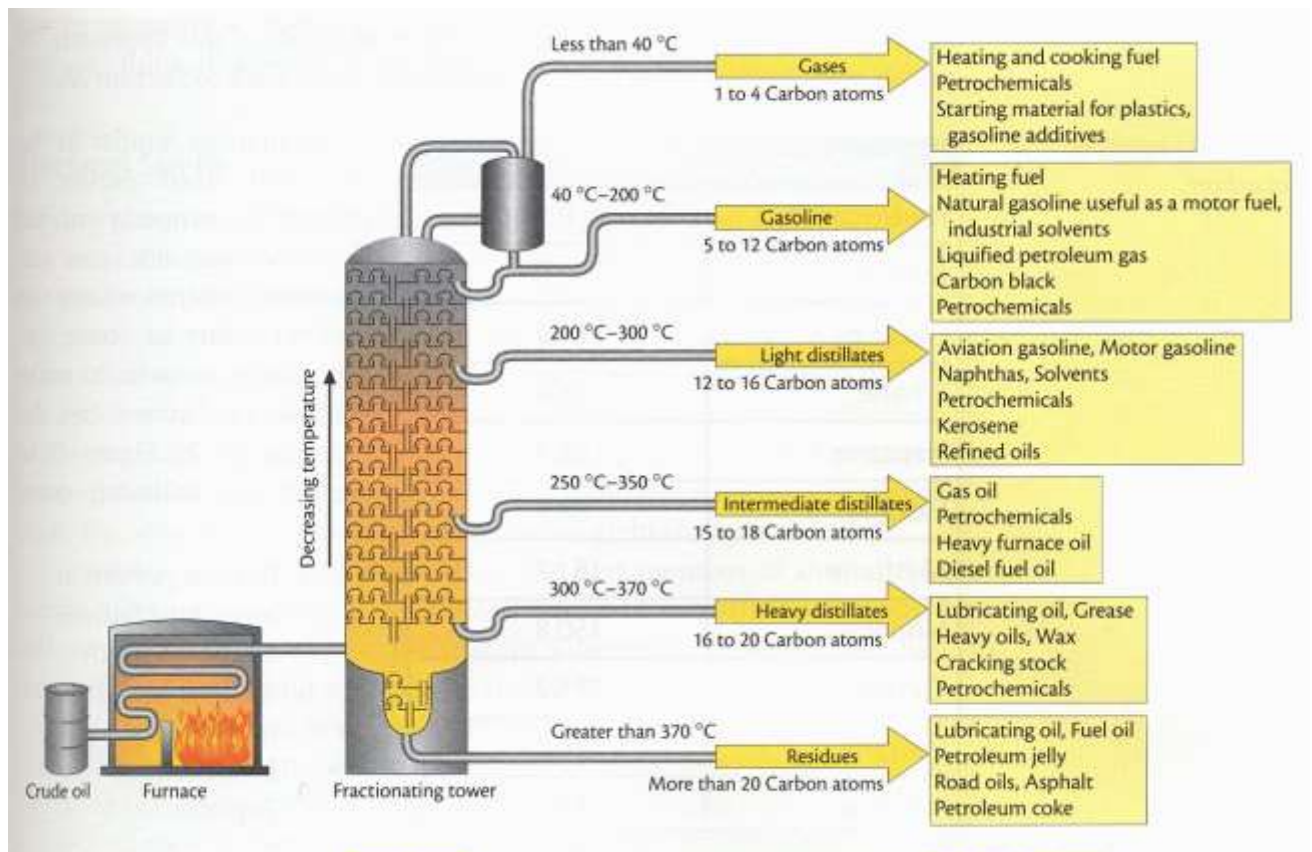
- Number of moles of iron started with?
 - Moles of tri Iron oxide formed
 - Mass of tri iron oxide formed
 - The concentration of this solution if we had 500ml of water in the reaction?
- 4) 25 ml of 0.1 M HCl reacts with 50ml of NaOH solution fully
What is the concentration of the NaOH solution.



Chapter 5

Organic chemistry

This is a branch of chemistry that looks at compounds of carbon chained molecules. The main source of these compounds is CRUDE OIL. We FRACTIONALLY DISTILL this to separate it out into different FRACTIONS which have similar boiling points, size and properties.



Task 15

Imagine you are a small CH_4 molecule in crude oil and you are being fractionally distilled,

What happened to you?

Why?

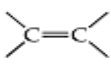
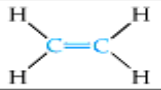
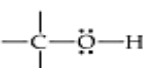
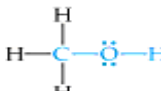
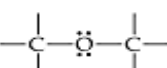
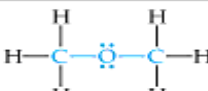
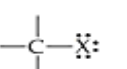
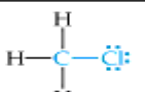
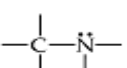
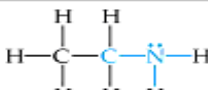
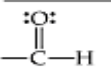
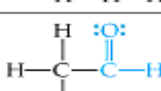

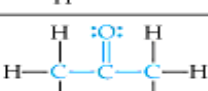
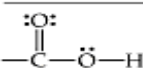
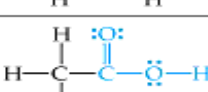
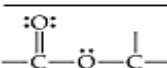
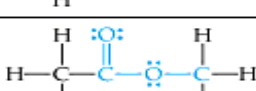
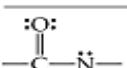
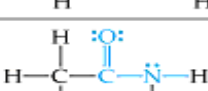
What happens to other molecules at the same time?

Why?

USE correct technical language to explain what's going on.

Types of organic compound

There are lots of different types of organic compound which are based upon their FUNCTIONAL GROUPS or parts of the compound which determine how they react.

TABLE 25.4 Common Functional Groups in Organic Compounds				
Functional Group	Type of Compound	Suffix or Prefix	Example	Systematic Name (common name)
	Alkene	<i>ene</i>		Ethene (Ethylene)
$\text{—C}\equiv\text{C—}$	Alkyne	<i>-yne</i>	$\text{H—C}\equiv\text{C—H}$	Ethyne (Acetylene)
	Alcohol	<i>-ol</i>		Methanol (Methyl alcohol)
	Ether	<i>ether</i>		Dimethyl ether
 (X = halogen)	Haloalkane	<i>halo-</i>		Chloromethane (Methyl chloride)
	Amine	<i>-amine</i>		Ethylamine
	Aldehyde	<i>-al</i>		Ethanal (Acetaldehyde)
	Ketone	<i>-one</i>		Propanone (Acetone)
	Carboxylic acid	<i>-oic acid</i>		Ethanoic acid (Acetic acid)
	Ester	<i>-oate</i>		Methyl ethanoate (Methyl acetate)
	Amide	<i>-amide</i>		Ethanamide (Acetamide)

The table shows the most common functional groups with examples and naming ideas.

Another aspect of organic compounds is the SERIES (called HOMOLOGOUS SERIES) you have of compounds which all have the same functional group.

These all increase by $\text{-CH}_2\text{-}$ each time and have a common pattern of naming linked to the number of carbons in the compound.

Name	Molecular formula	Full structural formula
Methane	CH_4	$\begin{array}{c} \text{H} \\ \\ \text{H} - \text{C} - \text{H} \\ \\ \text{H} \end{array}$
Ethane	C_2H_6	$\begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H} - \text{C} - \text{C} - \text{H} \\ \quad \\ \text{H} \quad \text{H} \end{array}$
Propane	C_3H_8	$\begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{H} \\ \quad \quad \\ \text{H} \quad \text{H} \quad \text{H} \end{array}$
Butane	C_4H_{10}	$\begin{array}{c} \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \\ \quad \quad \quad \\ \text{H} - \text{C} - \text{C} - \text{C} - \text{C} - \text{H} \\ \quad \quad \quad \\ \text{H} \quad \text{H} \quad \text{H} \quad \text{H} \end{array}$

Task 16

Research

What are the FIRST 10 stem names for organic compounds
using alcohols as an example write out the molecular formula for the first 10, draw out the full structural/ displayed formula for the first 10 and names then as well.
(HINT complete a table like one above but for the first 10 alcohols!)

Chapter 6

Calculations on efficiency of reactions.

There are two main methods that are used to look over the efficiency of chemical reactions.

1) Atom economy



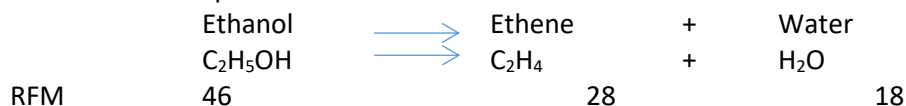
CALCULATION OF ATOM ECONOMY

$$\text{atom economy} = \frac{\text{mass of atoms in desired products}}{\text{mass of atoms in reactants}} \times 100\%$$

This is a measure of the useful products compared to all the products.

e.g.

Ethanol is decomposed into useful ethane and waste water.



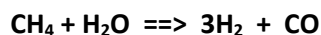
$$\begin{aligned}
 \text{Atom economy} &= \frac{\text{mass of useful product}}{\text{mass of all reactants}} \times 100 \\
 &= \frac{28}{46} \times 100 \\
 &= \underline{\underline{60.9\%}}
 \end{aligned}$$

Task 17

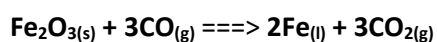
What is the Atom economy in:

1. Hydrogen is used in synthesising ammonia and is made on a large scale from reacting methane with water

methane + water ==> hydrogen + carbon monoxide



2. In the blast furnace where we form Iron .



2) Percentage yield

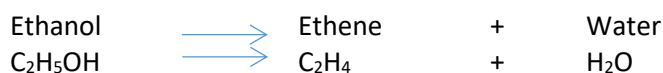
This is the second method we use to calculate the efficiency of the reaction. This gives an idea of what is actually formed in reality as compared to what we would expect to be formed.

$$\text{Percent Yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100\%$$

NOTE

You are often given the actual amount you form BUT you have to work out the theoretical amount from a mole calculation.

e.g. Ethanol is decomposed into useful ethane and waste water.



We create 1.4 g of the ethene from a starting mass of 4.6g of ethanol, what is the percentage yield.

CALC Moles = $\frac{\text{Mass}}{\text{RFM}}$

$$\text{Moles} = \frac{4.6}{46}$$

$$= 0.1 \text{ moles}$$

$$\begin{array}{l} 0.1 \text{ moles} : 0.1 \text{ moles} \\ \text{Mass} = \text{Moles} \times \text{RFM} \\ = 0.1 \times 28 \\ = 2.8 \text{ g} \end{array}$$

This is the theoretical yield amount

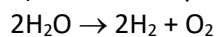
i.e this is the full amount that could possibly be formed

$$\begin{array}{lcl} \text{Final calc} & \text{percentage yield} & = \frac{\text{Actual}}{\text{Theoretical}} \times 100 \\ & & = \frac{1.4}{2.8} \times 100 \\ & & = 50\% \end{array}$$

Task 18

- 1) When 5.00 g of KClO_3 is heated it decomposes according to the equation: $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$
- Calculate the theoretical yield of oxygen.
 - Give the % yield if 1.78 g of O_2 is produced.
 - How much O_2 would be produced if the percentage yield was 78.5%?

- 2) The electrolysis of water forms H_2 and O_2 .



What is the % yield of O_2 if 12.3 g of O_2 is produced from the decomposition of 14.0 g H_2O ?

Appendix I

Common ions

Positive Ions (cations)		Negative Ions (anions)	
Name	Formula	Name	Formula
Hydrogen	H^+	Chloride	Cl^-
Sodium	Na^+	Bromide	Br^-
Silver	Ag^+	Fluoride	F^-
Potassium	K^+	Iodide	I^-
Lithium	Li^+	Hydroxide	OH^-
Ammonium	NH_4^+	Nitrate	NO_3^-
Barium	Ba^{2+}	Oxide	O^{2-}
Calcium	Ca^{2+}	Sulphide	S^{2-}
Copper(II)	Cu^{2+}	Sulphate	SO_4^{2-}
Magnesium	Mg^{2+}	Carbonate	CO_3^{2-}
Zinc	Zn^{2+}	Hydrogencarbonate	HCO_3^-
Lead	Pb^{2+}		
Iron(II)	Fe^{2+}		
Iron(III)	Fe^{3+}		
Aluminium	Al^{3+}		

Wider Reading

News/ Articles

[Magazine archive 2024 | New Scientist](#)

Useful Websites

www.rsc.org – Royal Society of Chemistry

Blogs

[Chemistry student blogs \(warwick.ac.uk\)](#)

Viewing and Listening

Ted talks:

[Andrew Parsons: The future is bright as chemistry comes to light: think chemically...naturally. | TED Talk](#)

Podcasts

Free podcasts, video lectures, reading recommendations o A whole range of resources from leading universities (Oxford, Yale)

YouTube

Has its own educational channel – EDU

Documentaries

[The 21 best science documentaries you should watch right now | New Scientist](#)

Online Courses

Futurelearn